

ENC AND STABILITY

(Honors Chemistry)

<p style="text-align: center;"><u>Lithium</u></p> <p>Group: <u>1A</u></p> <p>Valence Electrons: <u>1</u></p> <p>Principle Quantum Number: <u>2</u></p> <p>Shielding Electrons: <u>2</u></p> <p>Effective Nuclear Charge: <u>+1</u></p> <p>Charge of Ion: <u>1+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>	<p style="text-align: center;"><u>Beryllium</u></p> <p>Group: <u>2A</u></p> <p>Valence Electrons: <u>2</u></p> <p>Principle Quantum Number: <u>2</u></p> <p>Shielding Electrons: <u>2</u></p> <p>Effective Nuclear Charge: <u>+2</u></p> <p>Charge of Ion: <u>2+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>	<p style="text-align: center;"><u>Boron</u></p> <p>Group: <u>3A</u></p> <p>Valence Electrons: <u>3</u></p> <p>Principle Quantum Number: <u>2</u></p> <p>Shielding Electrons: <u>2</u></p> <p>Effective Nuclear Charge: <u>+3</u></p> <p>Charge of Ion: <u>3+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>
<p style="text-align: center;"><u>Sodium</u></p> <p>Group: <u>1A</u></p> <p>Valence Electrons: <u>1</u></p> <p>Principle Quantum Number: <u>3</u></p> <p>Shielding Electrons: <u>10</u></p> <p>Effective Nuclear Charge: <u>+1</u></p> <p>Charge of Ion: <u>1+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>	<p style="text-align: center;"><u>Magnesium</u></p> <p>Group: <u>2A</u></p> <p>Valence Electrons: <u>2</u></p> <p>Principle Quantum Number: <u>3</u></p> <p>Shielding Electrons: <u>10</u></p> <p>Effective Nuclear Charge: <u>+2</u></p> <p>Charge of Ion: <u>2+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>	<p style="text-align: center;"><u>Aluminum</u></p> <p>Group: <u>3A</u></p> <p>Valence Electrons: <u>3</u></p> <p>Principle Quantum Number: <u>3</u></p> <p>Shielding Electrons: <u>10</u></p> <p>Effective Nuclear Charge: <u>+3</u></p> <p>Charge of Ion: <u>3+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>
<p style="text-align: center;"><u>Potassium</u></p> <p>Group: <u>1A</u></p> <p>Valence Electrons: <u>1</u></p> <p>Principle Quantum Number: <u>4</u></p> <p>Shielding Electrons: <u>18</u></p> <p>Effective Nuclear Charge: <u>+1</u></p> <p>Charge of Ion: <u>1+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>	<p style="text-align: center;"><u>Calcium</u></p> <p>Group: <u>2A</u></p> <p>Valence Electrons: <u>2</u></p> <p>Principle Quantum Number: <u>4</u></p> <p>Shielding Electrons: <u>18</u></p> <p>Effective Nuclear Charge: <u>+2</u></p> <p>Charge of Ion: <u>2+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>	<p style="text-align: center;"><u>Gallium</u></p> <p>Group: <u>3A</u></p> <p>Valence Electrons: <u>3</u></p> <p>Principle Quantum Number: <u>4</u></p> <p>Shielding Electrons: <u>28</u></p> <p>Effective Nuclear Charge: <u>+3</u></p> <p>Charge of Ion: <u>3+</u></p> <p>Ionic Size: Increase or <u>Decrease</u></p>

<p style="text-align: center;"><u>Nitrogen</u></p> <p>Group: <u>5A</u></p> <p>Valence Electrons: <u>5</u></p> <p>Principle Quantum Number: <u>2</u></p> <p>Shielding Electrons: <u>2</u></p> <p>Effective Nuclear Charge: <u>+5</u></p> <p>Charge of Ion: <u>3⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>	<p style="text-align: center;"><u>Oxygen</u></p> <p>Group: <u>6A</u></p> <p>Valence Electrons: <u>6</u></p> <p>Principle Quantum Number: <u>2</u></p> <p>Shielding Electrons: <u>2</u></p> <p>Effective Nuclear Charge: <u>+6</u></p> <p>Charge of Ion: <u>2⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>	<p style="text-align: center;"><u>Flurine</u></p> <p>Group: <u>7A</u></p> <p>Valence Electrons: <u>7</u></p> <p>Principle Quantum Number: <u>2</u></p> <p>Shielding Electrons: <u>2</u></p> <p>Effective Nuclear Charge: <u>+7</u></p> <p>Charge of Ion: <u>1⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>
<p style="text-align: center;"><u>Phosphorus</u></p> <p>Group: <u>5A</u></p> <p>Valence Electrons: <u>5</u></p> <p>Principle Quantum Number: <u>3</u></p> <p>Shielding Electrons: <u>10</u></p> <p>Effective Nuclear Charge: <u>+5</u></p> <p>Charge of Ion: <u>3⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>	<p style="text-align: center;"><u>Sulfur</u></p> <p>Group: <u>6A</u></p> <p>Valence Electrons: <u>6</u></p> <p>Principle Quantum Number: <u>3</u></p> <p>Shielding Electrons: <u>10</u></p> <p>Effective Nuclear Charge: <u>+6</u></p> <p>Charge of Ion: <u>2⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>	<p style="text-align: center;"><u>Chlorine</u></p> <p>Group: <u>7A</u></p> <p>Valence Electrons: <u>7</u></p> <p>Principle Quantum Number: <u>3</u></p> <p>Shielding Electrons: <u>10</u></p> <p>Effective Nuclear Charge: <u>+7</u></p> <p>Charge of Ion: <u>1⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>
<p style="text-align: center;"><u>Arsenic</u></p> <p>Group: <u>5A</u></p> <p>Valence Electrons: <u>5</u></p> <p>Principle Quantum Number: <u>4</u></p> <p>Shielding Electrons: <u>28</u></p> <p>Effective Nuclear Charge: <u>+5</u></p> <p>Charge of Ion: <u>3⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>	<p style="text-align: center;"><u>Selenium</u></p> <p>Group: <u>6A</u></p> <p>Valence Electrons: <u>6</u></p> <p>Principle Quantum Number: <u>4</u></p> <p>Shielding Electrons: <u>28</u></p> <p>Effective Nuclear Charge: <u>+6</u></p> <p>Charge of Ion: <u>2⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>	<p style="text-align: center;"><u>Bromine</u></p> <p>Group: <u>7A</u></p> <p>Valence Electrons: <u>7</u></p> <p>Principle Quantum Number: <u>4</u></p> <p>Shielding Electrons: <u>28</u></p> <p>Effective Nuclear Charge: <u>+7</u></p> <p>Charge of Ion: <u>1⁻</u></p> <p>Ionic Size: <u>Increase</u> or Decrease</p>