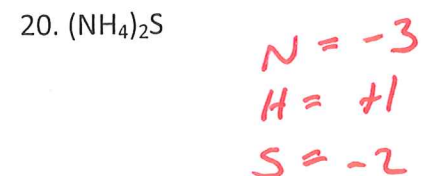
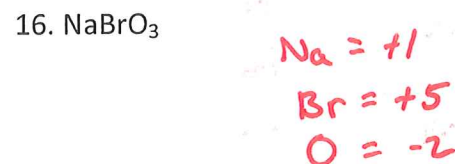
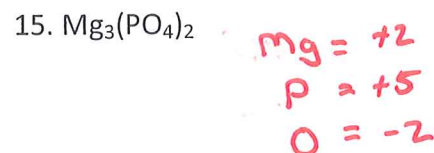
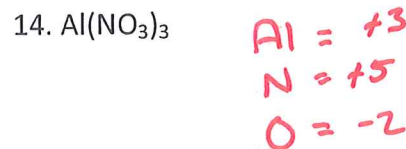
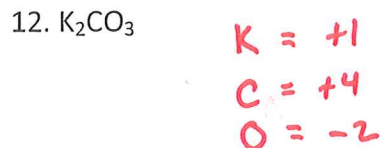
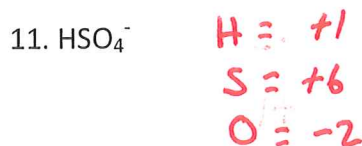
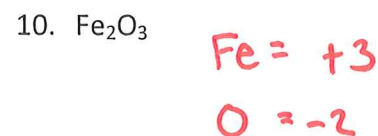
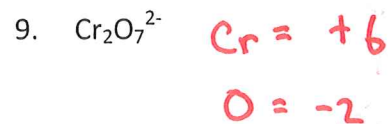
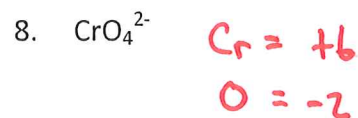
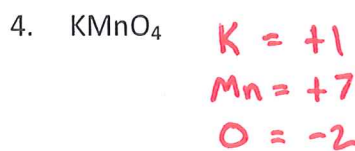
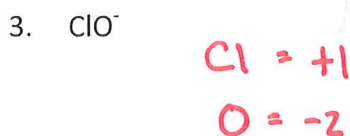
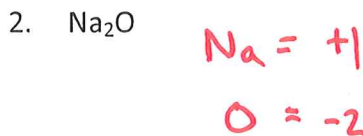


DETERMINING OXIDATION NUMBER

Honors Chemistry

* Calculate and list the oxidation numbers of each element in the following compounds.



Substance	Oxidation Number	Substance	Oxidation Number
$\begin{matrix} +2 & -2 \\ \text{Mg}(\text{ClO}_2)_2 \\ +2 & -8 \end{matrix}$	$\begin{matrix} \text{Mg} = +2 \\ \text{Cl} = +3 \\ \text{O} = -2 \end{matrix}$	H_2SO_3	$\begin{matrix} \text{H} = +1 \\ \text{S} = +4 \\ \text{O} = -2 \end{matrix}$
NH_4^+	$\begin{matrix} \text{N} = -3 \\ \text{H} = +1 \end{matrix}$	NaCl	$\begin{matrix} \text{Na} = +1 \\ \text{Cl} = -1 \end{matrix}$
PO_3^{3-}	$\begin{matrix} \text{P} = +3 \\ \text{O} = -2 \end{matrix}$	CN^-	$\begin{matrix} \text{C} = +2 \\ \text{N} = -3 \end{matrix}$
I_2	$\text{I} = 0$	Cl^-	$\text{Cl} = -1$
CO_2	$\begin{matrix} \text{C} = +4 \\ \text{O} = -2 \end{matrix}$	MnO_2	$\begin{matrix} \text{Mn} = +4 \\ \text{O} = -2 \end{matrix}$
HNO_2	$\begin{matrix} \text{H} = +1 \\ \text{N} = +3 \\ \text{O} = -2 \end{matrix}$	AgNO_3	$\begin{matrix} \text{Ag} = +1 \\ \text{N} = +5 \\ \text{O} = -2 \end{matrix}$
$\begin{matrix} +1 & -2 \\ \text{KClO}_3 \\ +1 & -6 \end{matrix}$	$\begin{matrix} \text{K} = +1 \\ \text{Cl} = +5 \\ \text{O} = -2 \end{matrix}$	$\text{Ca}(\text{OH})_2$	$\begin{matrix} \text{Ca} = +2 \\ \text{O} = -2 \\ \text{H} = +1 \end{matrix}$
$\begin{matrix} +1 & -2 \\ \text{Na}_2\text{Cr}_2\text{O}_7 \\ +2 & -14 \end{matrix}$	$\begin{matrix} \text{Na} = +1 \\ \text{Cr} = +6 \\ \text{O} = -2 \end{matrix}$	$\text{Ba}(\text{ClO}_3)_2$	$\begin{matrix} \text{Ba} = +2 \\ \text{Cl} = +5 \\ \text{O} = -2 \end{matrix}$
Cd^{2+}	$\text{Cd} = +2$	Cl_2	$\text{Cl} = 0$
FeCl_3	$\begin{matrix} \text{Fe} = +3 \\ \text{Cl} = -1 \end{matrix}$	H_2O_2	$\begin{matrix} \text{H} = +1 \\ \text{O} = -1 \end{matrix}$
SrO	$\begin{matrix} \text{Sr} = +2 \\ \text{O} = -2 \end{matrix}$	MgBr_2	$\begin{matrix} \text{Mg} = +2 \\ \text{Br} = -1 \end{matrix}$
$\begin{matrix} +5 & -2 \\ \text{AsO}_4^{3-} \\ -8 \end{matrix}$	$\begin{matrix} \text{As} = +5 \\ \text{O} = -2 \end{matrix}$	NH_3	$\begin{matrix} \text{N} = -3 \\ \text{H} = +1 \end{matrix}$
Fe^{3+}	$\text{Fe} = +3$	Ni	$\text{Ni} = 0$