

Reviewing Ionic Compounds

(Honors Chemistry)

* If the formula has a multi-valent metal, write both the stock and classical naming system!			
NaHCO ₃	Sodium hydrogen carbonate	Cr ₂ (Cr ₂ O ₇) ₃	Chromic dichromate Chromium III dichromate
Pb(OH) ₄	Plumbic hydroxide Lead IV hydroxide	Zn(OH) ₂	Zinc hydroxide
Pb ₃ P ₂	Plumbous phosphide Lead II phosphide	Ba(NO ₃) ₂	Barium nitrate
Fe ₂ (SO ₃) ₃	Ferric sulfite Iron III sulfite	Ca(MnO ₄) ₂	Calcium permanganate
FeCl ₂	Ferrous chloride Iron II chloride	NH ₄ NO ₂	Ammonium nitrite
AgF	Silver fluoride	KCl	Potassium chloride
Cu ₂ S	Cuprous sulfide Copper I sulfide	Cu(HCO ₃) ₂	Cupric hydrogen carbonate Copper II hydrogen carbonate
Sn(ClO) ₂	Stannous hypochlorite Tin II hypochlorite	Na ₃ N	Sodium nitride
Al(OH) ₃	Aluminum hydroxide	Zn(ClO) ₂	Zinc chlorite
Sn ₃ (PO ₄) ₂	Stannous phosphate Tin II phosphate	Ba(NO ₂) ₂	Barium nitrite
Cr ₂ (SO ₄) ₃	Chromic sulfate Chromium III sulfate	HgCl ₂	Mercuric chloride Mercury II chloride
Pb(ClO ₃) ₂	Plumbous chlorate Lead II chlorate	Cu ₂ SO ₄	Cuprous sulfate Copper I sulfate
Sn(NO ₃) ₄	Stannic nitrate Tin IV nitrate	HgClO ₄	Mercurous perchlorate Mercury I perchlorate
FePO ₃	Ferric phosphite Iron III phosphite	Cr ₂ (C ₂ O ₄) ₃	Chromic oxalate Chromium III oxalate
MnCrO ₄	Manganous chromate Manganese II chromate	SnCO ₃	Stannous carbonate Tin II carbonate
AuC ₂ H ₃ O ₂	Aurous acetate Gold I acetate	Ni(CN) ₃	Nickel III cyanide

Writing Chemical Formulas Review

Write the correct formula in the space below.			
$\text{Cu}_3(\text{PO}_4)_2$	Cupric Phosphite	$\text{Sn}(\text{ClO}_2)_4$	Stannic chlorite
Au_2S_3	Auric Sulfide	ZnSO_4	Zinc Sulfate
MnC_2O_4	Manganous oxalate	$\text{Cr}(\text{OH})_3$	Chromic hydroxide
$\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_3$	Ferric acetate	$\text{Co}(\text{CN})_3$	Cobaltic cyanide
$\text{Zn}_3(\text{PO}_3)_2$	Zinc phosphite	$\text{Pb}(\text{ClO}_4)_4$	Plumbic perchlorate
$\text{Au}_2\text{Cr}_2\text{O}_7$	Gold (I) dichromate	$\text{Fe}_3(\text{PO}_4)_2$	Iron (II) phosphate
NH_4NO_3	Ammonium nitrate	$\text{Sn}(\text{CO}_3)_2$	Tin (IV) carbonate
$\text{Cd}(\text{ClO})_2$	Cadmium hypochlorite	$\text{Pb}(\text{C}_2\text{O}_4)_2$	Plumbic oxalate
$\text{Hg}(\text{ClO}_3)_2$	Mercury (II) chlorate	NaOH	Sodium hydroxide
PbCrO_4	Plumbous chromate	HgS	Mercuric Sulfide
$\text{Ca}(\text{ClO}_4)_2$	Calcium perchlorate	Fe_3P_2	Ferrous Phosphide
$\text{Cr}_3(\text{PO}_4)_2$	Chromous phosphate	$\text{Pb}_3(\text{PO}_4)_2$	Lead (II) phosphate
Fe_2S_3	Ferric Sulfide	$\text{Co}(\text{ClO}_2)_3$	Cobalt (III) Chlorite
KMnO_4	Potassium permanganate	$\text{Ni}_2(\text{Cr}_2\text{O}_7)_3$	Nickle (III) dichromate
Hg_2O	Mercurous oxide	NH_4Cl	Ammonium chloride
Au_3PO_4	Gold (I) phoshate	Sr_3N_2	Strontium nitride
HgNO_2	Mercurous nitrite	NaHCO_3	Sodium hydrogen carbonate